**CHEM 1311- Stoichiometry, Average Atomic Mass, Empirical Formulas**

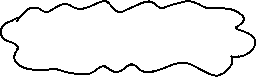
Calculate the number of formula units in 25.2 grams of sodium chloride.



Calculate the number of grams in 1.26x1023 atoms of water.



What is the percent mass of oxygen in copper (II) sulfate?



What is the percent mass of potassium in potassium hydroxide?



What is the empirical formula for the following compounds?

C4O12 \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_ Si2O4 \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_ N4H8Cl2 \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_



A compound contains 23.3% magnesium, 30.7% sulfur, and 46.0% oxygen. What is the empirical formula?



A compound contains 14.5% nitrogen, 61.3% carbon, and 24.2% sodium. What is the empirical formula?



Given that the empirical formula for cyclobutene is CH2 and it has a molar mass of 42g, what is its molecular formula?



Given that the empirical formula for benzene is CH and it has a molar mass of 78g, what is its molecular formula?

