**Metathesis Reactions 9/26**

List the strong acids:

List the strong bases:

Define a precipitation reaction and explain solubility in your definition.

Define an acid base reaction and include the net ionic product of all aqueous acid base reactions in your definition.

Write the ionic and net ionic equation for the following reactions:

Mg(NO3)2 (aq) + Na2CrO4 (aq) 🡪 MgCrO4 (s) + 2NaNO3 (aq)



Na3PO4 (aq) + FeCl3 (aq) 🡪 FePO4 (s) + 3NaCl (aq)





ZnCl2 (aq) + K2SO3 (aq) 🡪 2KCl (aq) + ZnSO3 (s)



