Chem 1311- Molarity, Dilutions, and Titrations

What is the molarity of a solution if 12.6 grams of NaCl dissolved in 0.280 liters of water?

How many milliliters of 5.0 M copper (II) sulfate solution must be added to dilute a 160 mL solution of 0.30M copper (II) sulfate?

Determine the molarity of a solution if 6.30 moles of KNO3 dissolved in 241 ml of water?

If you dilute 185 mL of a 2.50 M solution of LiCl to 2.00 L, determine the new concentration of the solution.

How many moles of solute exist if a 0.500M aqueous solution has a volume of 1.30 liters?

If 25.00 mL of HCl solution with a concentration of 0.750 M is neutralized by 23.45 mL of NaOH, what is the concentration of the base?

Titration reveals that 11.6 mL of 3.00 M sulfuric acid are required to neutralize the sodium hydroxide in 25.00 mL of NaOH solution. What is the molarity of the NaOH solution?

Create your own stoichiometric conversion road map: